Precipitation Net Ionics Practice

Write the double replacement reaction, the ionic reaction to determine spectators and then the net ionic reaction

A solution of copper II chloride is added to a solution of sodium sulfide

A solution of sodium hydroxide is added to a solution of lead (II) nitrate.

2007 a

A solution of barium chloride is added drop by drop to a solution of sodium carbonate, causing a precipitate to form. 2008B c

A barium nitrate solution and a potassium fluoride solution are combined and a precipitate forms. 2009B a

A 100 mL sample of 1 M strontium chloride solution is mixed with a 100 mL sample of

1 M sodium carbonate solution, resulting in the formation of a precipitate. 2011 B c

A solution of nickel (II) chloride is added to a solution of sodium hydroxide, forming a precipitate

Excess concentrated sulfuric acid is added to solid calcium phosphate

Hydrogen sulfide gas is bubbled into a solution of mercury II chloride

Solutions of manganese II sulfate and ammonium sulfide are mixed

An excess of sodium hydroxide solution is added to a solution of magnesium nitrate

A solution of copper II sulfate is added to a solution of barium hydroxide

Solid ammonium carbonate is added to a saturated solution of barium hydroxide

A solution of tri-potassium phosphate is added to a solution of zinc nitrate

Solutions of zinc sulfate and sodium phosphate are mixed

Solutions of silver nitrate and lithium bromide are mixed